

Lesson 20: Separation of Variables

II

Today's lecture we will be doing more of the same, plus some application problems.

Example 1: Find the general solution of the differential eqn:

$$(1) t^2 \frac{dy}{dt} - y = 0$$

Rewrite: $t^2 \frac{dy}{dt} = y$

$$\frac{dy}{y} = \frac{dt}{t^2}$$

$$\frac{dy}{y} = t^{-2} dt$$

$$\int \frac{dy}{y} = \int t^{-2} dt$$

$$\ln|y| = -t^{-1} + C$$

$$\ln|y| = -\frac{1}{t} + C$$

$$|y| = \exp\left[-\frac{1}{t} + C\right]$$

$$\pm y = e^C \exp\left[-\frac{1}{t}\right]$$

$$y = \underbrace{\pm e^C}_{\text{All a constant}} \exp\left[-\frac{1}{t}\right]$$

All a constant

$$y = C \exp\left[-\frac{1}{t}\right]$$

$$(2) -x^3 y + y' = 2x^3$$

Rewrite: $y' = 2x^3 + x^3 y$

$$\frac{dy}{dx} = x^3(2+y)$$

$$\frac{dy}{2+y} = x^3 dx$$

$$\int \frac{dy}{2+y} = \int x^3 dx$$

$$\ln|2+y| = \frac{x^4}{4} + C$$

$$|2+y| = \exp\left[\frac{x^4}{4} + C\right]$$

$$\pm(2+y) = e^C \exp\left[\frac{x^4}{4}\right]$$

$$2+y = \pm e^C \exp\left[\frac{x^4}{4}\right]$$

All a constant

$$2+y = C \exp\left[\frac{x^4}{4}\right]$$

$$y = C \exp\left[\frac{x^4}{4}\right] - 2$$

Example 2: In a particular chemical reaction, a substance is converted into a second substance at a rate proportional to the square of the amount of the first substance present at any time, t . Initially, 50 grams of the first substance was present, and 1 hour later only 14 grams of the first substance remained. What is the amount of the first substance remaining after 7 hours?

Set-Up: $\frac{da}{dt} = a^2 k$; $a(0) = 50$; $a(1) = 14$

Solve: $\frac{da}{a^2} = k dt$

$$\int a^{-2} da = \int k dt$$

$$-a^{-1} = kt + C$$

$$-\frac{1}{a} = kt + C$$

$$\frac{1}{a} = -kt - C$$

All a constant

$$\frac{1}{a} = -kt + C$$

$$a = \frac{1}{-kt + C}$$

When $a(1) = 14$,

$$14 = a(1) = \frac{50}{1 - 50k}$$

$$14(1 - 50k) = 50$$

$$1 - 50k = \frac{50}{14} = \frac{25}{7}$$

$$-50k = \frac{25}{7} - 1 = \frac{18}{7}$$

$$k = \frac{-1}{50} \cdot \frac{18}{7}$$

$$= -\frac{18}{350}$$

When $a(0) = 50$,
 $50 = a(0) = \frac{1}{C}$

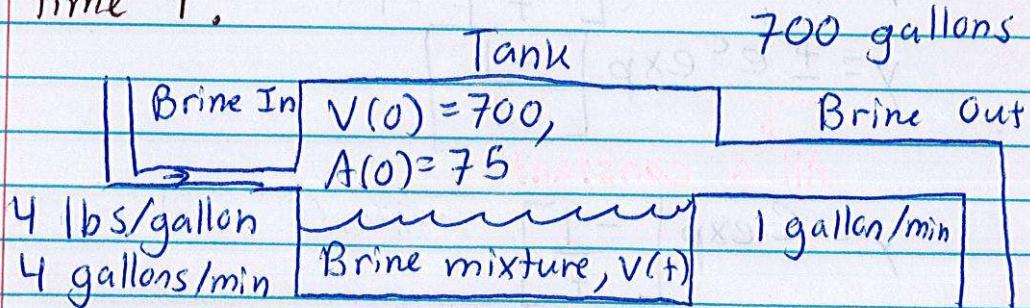
$$C = \frac{1}{50}$$

$$\begin{aligned} \text{So } a &= \frac{1}{\frac{1}{50} - kt} \\ &= \frac{50}{1 - 50kt} \end{aligned}$$

$$\begin{aligned} \text{So } a &= \frac{50}{1 - 50(-18) + 350} \\ &= \frac{350}{7 + 18t} \end{aligned}$$

Hence $a(7) \approx 2.6316$ grams

Example 3: A 800 gallon tank initially contains 700 gallons of brine containing 75 pounds of dissolved salt. Brine containing 4 pounds of salt per gallon flows into the tank at the rate of 4 gallons per minute and the well-stirred mixture flows out of the tank at the rate of 1 gallon per minute. Set up a differential equation for the amount of salt, $A(t)$, in the tank at time t .



Define: $V(t)$ = amount of brine mixture in tank at time t (in gallons)

$A(t)$ = amount of salt in the tank at time t
 t = time in minutes

$$\frac{dA}{dt} = (\text{rate of change of salt}) = (\text{rate in}) - (\text{rate out})$$

in tank in lbs/min of salt of salt

Rate in: $(4 \frac{\text{lbs}}{\text{gallons}})(4 \frac{\text{gallons}}{\text{min}}) = 16 \frac{\text{lbs}}{\text{min}}$

Rate out: "Well-stirred" means each gallon in the tank has as much salt in it as any other gallon.
i.e. Salt is uniformly mixed in the brine mixture

$$= \left(\frac{A(t)}{V(t)} \frac{\text{lbs}}{\text{gallons}} \right) \left(\frac{1 \text{ gallon}}{\text{min}} \right) = \frac{A(t)}{V(t)} \frac{\text{lbs}}{\text{min}}$$

$$\frac{dA}{dt} = 16 - \frac{A(t)}{V(t)}. \text{ Now find } V(t),$$

So $\frac{dV}{dt} = (\text{rate of change of brine mix}) = (\text{rate in}) - (\text{rate out})$

$$= \frac{4 \text{ gallons}}{\text{min}} - \frac{1 \text{ gallon}}{\text{min}} = \frac{3 \text{ gallons}}{\text{min}}$$

Hence $\begin{cases} V'(t) = 3 \\ V(0) = 700 \end{cases}$

$$\text{So } V(t) = \int V'(t) dt = \int 3 dt = 3t + C$$

$$\text{When } V(0) = 700,$$

$$700 = V(0) = 3(0) + C$$

$$700 = C$$

$$\Rightarrow V(t) = 3t + 700$$

Hence $\frac{dA}{dt} = 16 - \frac{A}{3t + 700}$